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Solid Acids and Bases Acidic and Basic Solvents The pH Scale Acid-Base Catalysis II **A Special Double Issue Devoted to Advances in Acidic and Basic Solid Materials Handbook of Reagents for Organic Synthesis, Acidic and Basic Reagents** The Chemistry of Non-aqueous Solvents: Acidic and basic solvents A Special Double Issue Devoted to Advances in Acidic and Basic Solid Materials **The Chemistry of Non-aqueous Solvents Acid-Base Diagrams Acidic and Basic Reagents** *Basic Concepts in Medicinal Chemistry Principles of Modern Chemistry* **Acids and Bases Acid-base Balance Superacids and Acidic Melts as Inorganic Chemical Reaction Media** A Physico-chemical Investigation of Solutions Containing Metallic Oxides (acidic and Basic) and Organic Hydroxy Acids *Methods for Characterizing the Acidic and Basic Surface Sites of Coal* **Influence of Acidic and Basic Properties on Selectivity of Reforming Reaction of N-hexane by Titania-Zirconia Catalyst Supporting Platinum** An Introduction to Chemistry **A System of Qualitative Chemical Analysis for the Common Basic and Acidic Ions** Determination of Acidic and Basic Species in Particulates by Thermometric Titration Calorimetry *Reactions of Acids and Bases in Analytical Chemistry* **Characterization of Acidic and Basic Properties of Heterogeneous Catalysts by Test Reactions Chemistry 2e** *Production and Regulation of Acidic and Basic Fibroblast Growth Factors in the Mammary Gland* **Acidic Proteins of the Nucleus** *Salt, Fat, Acid, Heat* Porphyrins in Super Acidic, Acidic, and Basic Media Qpnc-Page *Certain Acidic and Basic Catalyzed Condensations and Cyclizations ...* **Anatomy & Physiology** *Polymer Composite Materials — Interface Phenomena & Processes* Physico-chemical Investigation of Acidic and Basic Gaseous and Particulate Pollutants in the Troposphere Calculating Exact Charge & pI of Amino acids, Peptides and other Molecules **A Preliminary Study of the Interaction of Acidic and Basic Drugs Using Ethyl Cellulose Microspheres Acid-base Cements** *Chemistry Oxygen Soft Mechanochemical Synthesis* **Researches, Chemical and Philosophical; Chiefly Concerning Nitrous Oxide**

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The pH scale measures how acidic or basic a substance is, ranging from 0 to 14. Readers will learn how certain substances rank on the pH scale, what happens when acids and bases are mixed, and how water can make a substance either acidic or basic. These significant science concepts are discussed in approachable text and supported by motivating fact boxes, charts, images, and photographs. This book teaches chemistry at an appropriate level of rigor while removing the confusion and insecurity that impair student success. Students are frequently intimidated by prep chem; Bishop's text shows them how to break the material down and master it. The flexible order of topics allows unit conversions to be covered either early in the course (as is traditionally done) or later, allowing for a much earlier than usual description of elements, compounds, and chemical reactions. The text and superb illustrations provide a solid conceptual framework and address misconceptions. The book helps students to develop strategies for working problems in a series of logical steps. The Examples and Exercises give plenty of confidence-building practice; the end-of-chapter problems test the student's mastery. The system of objectives tells the students exactly what they must learn in each chapter and where to find it. Acidic Proteins of the Nucleus focuses on the functional role of acidic nuclear proteins in differential gene expression. Historically, these proteins are referred to as acidic in nature because they are insoluble in dilute mineral acids and their amino acid composition shows a preponderance of acidic over basic amino acid residues. After an introduction to DNA-binding proteins and transcriptional control in prokaryotic and eukaryotic systems, the subsequent chapters describe various approaches for isolating, separating, and characterizing acidic nuclear proteins. The core chapters specifically cover the isolation, fractionation, and characterization of acidic nuclear phosphoproteins, and the role of these proteins in cell proliferation, cell differentiation, and cell cycle. The last two chapters address the role of acidic nuclear protein in binding steroid hormones and in gene regulation. Each chapter contains some previously unpublished work and provides recommendations for future research. This book will be a good reference background for researchers of acidic nuclear proteins. Research Paper from the year 2012 in the subject Medicine - Neurology, Psychiatry, Addiction, language: English, abstract: QPNC-PAGE (abbreviation for: quantitative preparative native continuous polyacrylamide gel electrophoresis) is a standardized variant of the electrophoresis, particularly gel electrophoresis. This analytical method of biochemistry and bioinorganic chemistry is used for the separation of charged molecules in a homogeneous electric field and allows the quantitative separation and isolation of metalloproteins from human, vegetable or animal samples with high resolution. Proteins with different molecular mass and charge are separated according to isoelelectric points and analyzed by nuclear magnetic resonance (NMR). Thus the method makes an important contribution to the structure determination of native and denatured metalloproteins and protein isomers in complex protein mixtures. New technologies demand new materials. Polymer composites, with their wide range of possible fillers and polymers, open the way to an enormous range of materials with differing chemical, physical, and mechanical properties. The ultimate goal of polymer composite research is to formulate procedures that will lead to the design of composites with preset, i.e. specified, properties. Based on many years' experience in the field, the authors prepare the way towards just such a design procedure. The key element is the analysis and classification of the state of the filler-polymer interfaces from the point of view of their acid-base adsorption interactions. These interfacial phenomena play a pivotal role in determining overall properties of the composite: its rheological behaviour, its structural properties, catalytic effects in polymerization and polycondensation, and other technological characteristics. The book discusses and evaluates the extensive previous research scattered throughout the literature in Eastern Europe and the West, presents numerous experimental studies, and sets new benchmarks for the analysis of polymer composites. The book is required for researchers wanting to keep abreast of the progress in the burgeoning fields of polymer analysis and design. Acids and bases are ubiquitous in chemistry. Our understanding of them, however, is dominated by their behaviour in water. Transfer to non-aqueous solvents leads to profound changes in acid-base strengths and to the rates and equilibria of many processes: for example, synthetic reactions involving acids, bases and nucleophiles; isolation of pharmaceutical actives through salt formation; formation of zwitter- ions in amino acids; and chromatographic separation of substrates. This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non-aqueous solvents. The behaviour is related where possible to that in water, but correlations and contrasts between solvents are also presented. Fundamental background material is provided in the initial chapters: quantitative aspects of acid-base equilibria, including definitions and relationships between solution pH and species distribution; the influence of molecular structure on acid strengths; and acidity in aqueous solution. Solvent properties are reviewed, along with the magnitude of the interaction energies of solvent molecules with (especially) ions; the ability of solvents to participate in

hydrogen bonding and to accept or donate electron pairs is seen to be crucial. Experimental methods for determining dissociation constants are described in detail. In the remaining chapters, dissociation constants of a wide range of acids in three distinct classes of solvents are discussed: protic solvents, such as alcohols, which are strong hydrogen-bond donors; basic, polar aprotic solvents, such as dimethylformamide; and low-basicity and low polarity solvents, such as acetonitrile and tetrahydrofuran. Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents, and there is a strong differentiation between the response of neutral and charged acids to solvent change. Ion-pairing and hydrogen-bonding equilibria, such as between phenol and phenoxide ions, play an increasingly important role as the solvent polarity decreases, and their influence on acid-base equilibria and salt formation is described. This study was done to enhance the compatibility between an acidic and basic drug when used in combination within the same dosage form by encapsulating the drugs within an ethyl cellulose microsphere. The drugs encapsulated include acidic (acetaminophen and aspirin) and basic (theophylline, and hydroxyzine) drugs. Microspheres were prepared by the solvent evaporation technique using methylene chloride and acetone as the dispersed phase and light mineral oil with span 80 as the continuous phase. Various drug to polymer ratios (1:1, 1:2, and 1:3) were chosen and evaluated to determine which microspheres resulted in the highest drug entrapment efficiency. The microspheres with highest drug entrapment were characterized by performing particle size analysis, differential scanning calorimetry (DSC), scanning electron microscopy (SEM), fourier transform infrared spectroscopy (FTIR), and in vitro drug release. Stability studies were performed with a (1:1) mixture of an encapsulated acidic and basic drug at 40 °C for three months. The 1:1 ratio had the highest entrapment for the drugs aspirin (32%) and hydroxyzine (41%), while the 1:3 ratio had the highest entrapment for acetaminophen (65%) and theophylline (52%). Spherical, uniform microspheres were collected and the average particle size was found for acetaminophen (91.17 ?m), aspirin (193.36 ?m), hydroxyzine (93.89 ?m) and theophylline (264.56 ?m). DSC thermograms for the drug loaded microspheres determined that the drug present exists in an amorphous phase in the polymeric matrix indicating no interaction between the drug and the polymer. The in vitro drug release studies showed that acetaminophen and hydroxyzine released in 15-30 minutes and aspirin and theophylline released in 180-240 minutes. The drug release data was fitted to various kinetic models which resulted in acetaminophen following first order release (R2 value of 0.9693), aspirin following zero order release (R2 value of 0.9925), hydroxyzine following first order release (R2 value of 0.9744) and theophylline following Korsemeyer-Peppas release (R2 value of 0.9296). FTIR results, over three months, showed inconclusive results, so further analysis is required. After evaluation of the microspheres, the data collected indicates the microspheres, when used in combination, need to be further analyzed using HPLC to determine if any interactions are occurring between the acidic and basic drugs. Solid acid catalysts are already being used in various processes in petroleum refining and are presently being studied intensively in both academic and applied fields for usage in a variety of reactions. Solid base catalysts are also gaining increasing recognition as potential catalysts. Both acidic and basic catalysts are promising not only with respect to acid and base-catalyzed reactions but also in materials sciences, such as the production of adsorbents, sensors, ceramics, etc. The present volume presents the text of 21 invited oral presentations and 58 poster presentations. The material covers a wide range of aspects on acid-base catalysis, from quantum chemistry to industrialized processes. Medicinal chemistry is a complex topic. Written in an easy to follow and conversational style, Basic Concepts in Medicinal Chemistry focuses on the fundamental concepts that govern the discipline of medicinal chemistry as well as how and why these concepts are essential to therapeutic decisions. The book emphasizes functional group analysis and the basics of drug structure evaluation. In a systematic fashion, learn how to identify and evaluate the functional groups that comprise the structure of a drug molecule and their influences on solubility, absorption, acid/base character, binding interactions, and stereochemical orientation. Relevant Phase I and Phase II metabolic transformations are also discussed for each functional group. Key features include: • Discussions on the roles and characteristics of organic functional groups, including the identification of acidic and basic functional groups. • How to solve problems involving pH, pKa, and ionization; salts and solubility; drug binding interactions; stereochemistry; and drug metabolism. • Numerous examples and expanded discussions for complex concepts. • Therapeutic examples that link the importance of medicinal chemistry to pharmacy and healthcare practice. • An overview of structure activity relationships (SARs) and concepts that govern drug design. • Review questions and practice problems at the end of each chapter that allow readers to test their understanding, with the answers provided in an appendix. Whether you are just starting your education toward a career in a healthcare field or need to brush up on your organic chemistry concepts, this book is here to help you navigate medicinal chemistry. About the Authors Marc W. Harrold, BS, Pharm, PhD, is Professor of Medicinal Chemistry at the Mylan School of Pharmacy, Duquesne University, Pittsburgh, PA. Professor Harrold is the 2011 winner of the Omicron Delta Kappa "Teacher of the Year" award at Duquesne University. He is also the two-time winner of the "TOPS" (Teacher of the Pharmacy School) award at the Mylan School of Pharmacy. Robin M. Zavod, PhD, is Associate Professor for Pharmaceutical Sciences at the Chicago College of Pharmacy, Midwestern University, Downers Grove, IL, where she was awarded the 2012 Outstanding Faculty of the Year award. Professor Zavod also serves on the adjunct faculty for Elmhurst College and the Illinois Institute of Technology. She currently serves as Editor-in-Chief of the journal Currents in Pharmacy Teaching and Learning. Solid Acids and Bases: Their Catalytic Properties reviews developments in the studies of acidic and basic properties of solids, including the efficacy and special characteristics of solid acid and base catalysts. This book discusses the determination of basic and acidic properties on solid surfaces and relationship between acid strength and acid amount. The structure and acid-base properties of mixed metal oxides and correlation between acid-base properties and catalytic activity and selectivity are also deliberated. This publication is useful to professional chemists and graduate students in the fields of organic, inorganic and physical chemistry, petroleum chemistry and catalysis, including readers interested in the acidic and basic properties on solid

surfaces. "Davy discovered the anaesthetic properties of nitrous oxide and suggested its use during surgical operations ..."--Garrison-Morton. Now a Netflix series New York Times Bestseller and Winner of the 2018 James Beard Award for Best General Cookbook and multiple IACP Cookbook Awards Named one of the Best Books of 2017 by: NPR, BuzzFeed, The Atlantic, The Washington Post, Chicago Tribune, Rachel Ray Every Day, San Francisco Chronicle, Vice Munchies, Elle.com, Glamour, Eater, Newsday, Minneapolis Star Tribune, The Seattle Times, Tampa Bay Times, Tasting Table, Modern Farmer, Publishers Weekly, and more. A visionary new master class in cooking that distills decades of professional experience into just four simple elements, from the woman declared "America's next great cooking teacher" by Alice Waters. In the tradition of The Joy of Cooking and How to Cook Everything comes Salt, Fat, Acid, Heat, an ambitious new approach to cooking by a major new culinary voice. Chef and writer Samin Nosrat has taught everyone from professional chefs to middle school kids to author Michael Pollan to cook using her revolutionary, yet simple, philosophy. Master the use of just four elements--Salt, which enhances flavor; Fat, which delivers flavor and generates texture; Acid, which balances flavor; and Heat, which ultimately determines the texture of food--and anything you cook will be delicious. By explaining the hows and whys of good cooking, Salt, Fat, Acid, Heat will teach and inspire a new generation of cooks how to confidently make better decisions in the kitchen and cook delicious meals with any ingredients, anywhere, at any time. Echoing Samin's own journey from culinary novice to award-winning chef, Salt, Fat Acid, Heat immediately bridges the gap between home and professional kitchens. With charming narrative, illustrated walkthroughs, and a lighthearted approach to kitchen science, Samin demystifies the four elements of good cooking for everyone. Refer to the canon of 100 essential recipes--and dozens of variations--to put the lessons into practice and make bright, balanced vinaigrettes, perfectly caramelized roast vegetables, tender braised meats, and light, flaky pastry doughs. Featuring 150 illustrations and infographics that reveal an atlas to the world of flavor by renowned illustrator Wendy MacNaughton, Salt, Fat, Acid, Heat will be your compass in the kitchen. Destined to be a classic, it just might be the last cookbook you'll ever need. With a foreword by Michael Pollan. By doing both pH and thermometric titrations with NaOH and HCl on water extracts of particulate samples it is usually possible to positively identify the kinds and amounts of acidic and basic species present in the sample. The acid ionization constant of an unknown acid or base can be obtained from the pH titration curve. The heat of proton ionization, .delta.H/sub i/, from the same unknown can be obtained from the thermogram. These two constants (pK, .delta.H/sub i/) are usually sufficient to identify the unknown. Results of such studies of samples of dust from smelters and of an urban aerosol sample are used to illustrate the method. The principal acidic and basic species in the dust samples were metal ions (Ca(II), Pb(II), Zn(II), Cu(II), Fe(II), Fe(III)) which were largely identifiable from the data. The urban aerosol was found to contain carboxylate ions, ammonium ions, phenol, and primary alkyl ammonium ions. Acid sulfate species such as HSO/sup -//sub 4/ and H/sub 2/SO/sub 4/ were shown not to be present in the extractant, although SO/sup 2 -//sub 4/ was present. A titania-zirconia catalyst supporting platinum, which has both acid and base sites, was applied to a catalytic reforming reaction of n-hexane. The relationships were investigated between acid and base amounts of the catalyst, and rates of the cracking, isomerization, cyclic and aromatization reactions. The acid and base amounts were varied by changing the composition of titania in the catalyst. The rates of cyclic and aromatization reactions linearly increased with the amount of acid sites which make pairs with adjacent base sites (refered to adjacent pairs of acid-base sites), and the rates of cracking and isomerization reactions were proportional to the amount of the remaining acid sites (lone acid sites). Classification of catalysts according to their active sites can be done by catalytic test reactions. In contrast to acid catalysed test reactions which are actually well understood there is a need to study base catalysed reactions in detail. Therefore reactions in gas phase like conversion of methyl butynol (MBOH) and isopropanol and in liquid phase like Knoevenagel condensation were investigated. It was found that the conversion of isopropanol yields propene as the only product over the catalysts which have no redox ability while the reaction was sensitive to nature of the active centers in case of redox active catalysts. For the conversion of MBOH, the formation of MBYNE was an indication for acidic sites in the investigated solids whereas acetone and acetylene were found to be the products of the basic pathway. In addition, a new mechanism was proposed for the formation of the product 3-methyl-3-butyn-2-on (MIPK) requiring not only strong acid sites but also a special structural environment. A comparison between MBOH conversion and Knoevenagel condensation shows that the two test reactions used in the present study characterize the same basic properties proving the explanatory power of the test reaction. engl. Mechanical methods of the activation of chemical processes are currently widely used for the synthesis of various compounds. The present monograph deals with the development of a novel approach to mechanochemical synthesis based on reactions of solid acids, bases, hydrated compounds, crystal hydrates, basic and acidic salts. This method has been called soft mechanochemical synthesis. The monograph includes the papers published by the present authors. They describe the results of their investigations n the last two decades. New theoretical and experimental data on kinetics and mechanism of soft mechanochemical reactions in the mixtures of compounds mentioned above to give complex oxide compounds are presented. The description of new high energetic and high efficient mills providing effective occurrence of these reactions is delivered. The possibilities of applying soft mechanochemical synthesis for materials used in catalysts, material science, electronics, etc., are discussed. The advantages of the method proposed in comparison with other methods are demonstrated. The monograph is designed for researchers, engineers and technicians engaged in chemical and ceramic industry, for scientists and students specialized in the area of development, and application of new materials. This book is the first comprehensive account of acid-base reaction cements. These materials, which are formed by reacting an acid and a base, offer an alternative to polymerisation as a means of forming solid substances. Understanding acid-base equilibria made easy for students in chemistry, biochemistry, biology, environmental and earth sciences. Solving chemical problems, be it in education or in real life, often requires the understanding of the acid-base equilibria behind them. Based on many

years of teaching experience, Heike Kahlert and Fritz Scholz present a powerful tool to meet such challenges. They provide a simple guide to the fundamentals and applications of acid-base diagrams, avoiding complex mathematics. This textbook is richly illustrated and has full color throughout. It offers learning features such as boxed results and a collection of formulae. Handbook of Reagents for Organic Synthesis Acidic and Basic Reagents Hans J. Reich University of Wisconsin at Madison, USA James H. Rigby Wayne State University, Detroit, USA Recognising the critical need for bringing a handy reference work that deals with the most popular reagents in synthesis to the laboratory of practising organic chemists, the Editors of the acclaimed Encyclopedia of Reagents for Organic Synthesis (EROS) have selected the most important and useful reagents employed in contemporary organic synthesis. Handbook of Reagents for Organic Synthesis: Acidic and Basic Reagents, presents a selection of articles on the most fundamental and versatile reagents for effecting organic transformations that were originally included in EROS. In selecting candidate entries for inclusion in this particular collection, the editors adopted a broad set of criteria for defining what exactly constitutes an acidic or basic reagent. Each article contains all of the information found in EROS as well as expanded related reagents listings. Additional new listings of recently published review articles and monographs are included, as well as relevant Organic Syntheses procedures that deal with either the preparations or reactions of the featured reagents. This thorough and comprehensive handbook will prove of widespread appeal. The book presents the mathematical aspects of calculating the exact charge and isoelectric pH (pI) of amino acids, peptides and other molecules including drugs and pH indicators. The methods presented in this textbook are derived from the classical Henderson-Hasselbalch equation for weak acids and bases. They can be applied to calculate the exact charge and pI of amino acids and peptides, and percentage and fraction abundance of the uncharged, ionized and zwitterion forms of the amino acid at any specified pH. The use of Excel or similar data processing software is recommended while dealing with peptides and proteins. The methods can be extended to several applications like calculation of charge and ionization of drugs and pH indicators molecules, etc. It thus enables the user to quantify charge and ionization of any molecule bearing weakly acidic and basic groups, and subsequently apply it as needed in many fields, from the classrooms to research laboratories. PRINCIPLES OF MODERN CHEMISTRY has dominated the honors and high mainstream general chemistry courses and is considered the standard for the course. The fifth edition is a substantial revision that maintains the rigor of previous editions but reflects the exciting modern developments taking place in chemistry today. Authors David W. Oxtoby and H. P. Gillis provide a unique approach to learning chemical principles that emphasizes the total scientific process'from observation to application'placing general chemistry into a complete perspective for serious-minded science and engineering students. Chemical principles are illustrated by the use of modern materials, comparable to equipment found in the scientific industry. Students are therefore exposed to chemistry and its applications beyond the classroom. This text is perfect for those instructors who are looking for a more advanced general chemistry textbook.

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